

Name \_\_\_\_\_

Date \_\_\_\_\_

Period \_\_\_\_\_

**Acids And Bases Ws #6: More Fun Practice**

	Solute	Name of solute	Molar Mass (g/ mole)	Molarity (mol/L)	grams solute Liter solution	[H <sup>+</sup> ]	[OH <sup>-</sup> ]	pH	pOH	acid or base
1.	HCl	Hydrochloric Acid	36.5	0.020	0.73	$2.0 \times 10^{-2}$	$5.0 \times 10^{-13}$	1.70	12.3	Acid
2.	NaOH			0.040						
3.	HClO <sub>4</sub>					$1.25 \times 10^{-3}$				
4.	HNO <sub>3</sub>				0.060					
5.	RbOH						$3.75 \times 10^{-2}$			
6.	H <sub>2</sub> S			0.0030						
7.	HNO <sub>3</sub>					$3 \times 10^{-4}$				
8.	KOH						$4.5 \times 10^{-2}$			
9.	Sr(OH) <sub>2</sub>							9.13		