

Name _____ Date _____ Period _____

Bonding Worksheet #6 More About Bonding

1. Describe how a covalent bond forms between two atoms.
2. Why is the diatomic molecule H_2 more stable than two separate hydrogen atoms?
3. Why do ionic compounds tend to have higher boiling and melting points than molecular compounds?
4. How does a covalent bond differ from an ionic bond?
5. Explain the difference between single, double, and triple bonds.
6. What is the name for the outermost energy level of an atom that can participate in bonding?
7. A structure in which atomic symbols represent nuclei and inner-shell electrons and in which dots are used to represent valence electrons is called a(n) _____.
8. The rule that says that atoms in a compound tend to have the electron configuration of a noble gas is called _____.
9. Methanol, CH_3OH , can be used as a solvent, as an antifreeze, and in the production of formaldehyde. Draw the Lewis structure for methanol.
10. Draw the Lewis structure for the rocket propellant nitryl flouride, NO_2F .