Name_____ Date_____ Period_____

Chemical Reactions Ws #3: Identifying Reaction Type

Directions: Write a balanced equation and indicate the reaction type for each of the following equations. SR = single replacement DR = double replacement
C = Combination D = Decomposition

_____ 1. Aluminum nitrate solution reacts with sodium hydroxide solution to yield sodium nitrate solution and a precipitate of aluminum hydroxide

_____ 2. Solid potassium chlorate is heated to yield solid potassium chloride and oxygen gas.

_____ 3. A solution of phosphoric acid reacts with magnesium hydroxide solution to give a precipitate of magnesium phosphate and liquid water.

_____ 4. In the presence of heat, ammonium nitrite solid will yield nitrogen gas and liquid water.

_____ 5. Solid Aluminum Sulfate breaks down to produce solid Aluminum Oxide and Sulfur trioxide gas.

_____ 6. Aqueous barium chloride reacts with aqueous sodium sulfate to give sodium chloride solution and a precipitate of barium sulfate.

_____ 7 A solution of iron III bromide reacts with aqueous ammonium sulfide to give solid iron III sulfide and aqueous ammonium bromide.

_____ 8. Solid calcium oxide combines with solid diphosphourus pentoxide to produce solid calcium phosphate

_____ 9. An aqueous solution of magnesium chloride combines with an aqueous solution of silver nitrate to produce aqueous magnesium nitrate and solid silver chloride.

_____ 10. Solid aluminum hydroxide and acetic acid solution combine to produce aqueous aluminum acetate and liquid water.

_____ 11. Solid iron reacts with solid silver acetate to yield iron II acetate solution plus beautiful solid silver.

_____ 12 Calcium hydroxide solution reacts with phosphoric acid to yield the salt precipitate calcium phosphate and water.

_____ 13. Sodium hydroxide solution reacts with sulfuric acid to give the soluble salt sodium sulfate and water

_____ 14 Solutions of Calcium Nitrate and Sulfuric acid are mixed and produce a solid precipitate of Calcium sulfate in a nitric acid solution.

_____ 15. A solution of Sulfuric Acid is synthesized from liquid water and sulfur trioxide gas.