

Name_____

Date_____

Period_____

Chemical Reactions Ws #5: Double Replacement Reactions

Use the "Solubility of Common Compounds" reference sheet to complete the following.

1. When the following solutions are mixed together, what precipitate (if any) will form?
 - a) $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq})$
 - b) $\text{Pb}(\text{NO}_3)_2(\text{aq}) + \text{KCl}(\text{aq})$
 - c) $\text{AgNO}_3(\text{aq}) + \text{Na}_3\text{PO}_4(\text{aq})$
 - d) $\text{NaOH} + \text{Fe}(\text{NO}_3)_3$
2. For the reactions in 1, write the balanced equation for the possible double replacement reaction. If no precipitate forms, write "No Reaction."
 - a)
 - b)
 - c)
 - d)
3. For the reactions in 1, write a complete ionic equation.
 - a)
 - b)
 - c)
 - d)
4. For the reactions in 1, write a net ionic equation.
 - a)
 - b)
 - c)
 - d)
5. Write the balanced equation and the net ionic equation for each of the following solution mixtures. If no precipitate forms, write "No Rxn"
5. Ammonium Sulfate and Barium Nitrate
6. Lead(II) Nitrate and sodium chloride
7. Sodium Phosphate and Potassium nitrate
8. Sodium bromide and rubidium chloride
9. Copper(II) chloride and Sodium Hydroxide